**Atomic Structure**

**Definitions**

Atom – The smallest particle of an element that is still recognisable as the element

Molecule – A group of atoms chemically joined together

**Sub-atomic particles**

You need to be able to describe the location, mass and charge of each particle

|  |  |  |  |
| --- | --- | --- | --- |
| Particle | Location | Mass | Charge |
|  |  |  |  |
|  |  |  |  |
|  |  |  |  |

**More definitions**

Atomic number –

Mass number –

Isotopes –

Relative atomic mass – Average mass of an atom of an element compared with an atom of carbon-12

**Calculating RAM**

RAM = (% abundance x mass) + (% abundance x mass)

 100

Mg24 = 79%

Mg25 = 11%

Mg26 = 10%

