**Born Haber – Definitions**

*Lattice Enthalpy ∆LEH*

The enthalpy change when 1 mole of solid ionic lattice is formed from its ions in their gaseous state. (Sometimes you will see this described as the reverse reaction)

*Enthalpy Change of Formation ∆fH*

The enthalpy change when 1 mole of substance is formed from it elements in their standard state

*Enthalpy Change of Atomisation ∆atH*

The enthalpy change when 1 mole of gaseous atoms are formed from the element in its standard state

*First Ionisation Energy I.E.*

The energy required to remove 1 mole of electrons from 1 mole of atoms in the gaseous state

*Electron Affinity E.A.*

The enthalpy change when 1 mole of electrons is gained by 1 mole of atoms in the gaseous state

